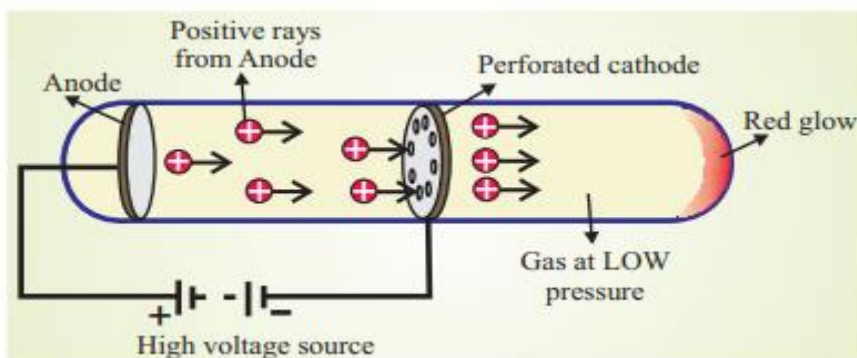


National Institute of Open Schooling
Secondary Course : Science And Technology
Lesson 5 : Atomic Structure
Worksheet -5

1. Look at the figure carefully:



- a. What does this figure show?
 - b. Who has done this experiment?
 - c. Explain this experiment in detail.
2. Write the electronic configuration of the elements having 12,14, 19, 23 and 26 protons in their nucleus.
 3. Explain the plum pudding model of the atom.
 4. In what way are the different atomic particles arranged in the atom? Explain it with reasons.
 5. A helium atom has 2 electrons in its outermost shell but its valency is 0. Explain
 6. "The number of protons and electrons in an atom and its corresponding ions are different." Justify this statement by using knowledge of the atomic structure.
 7. If the atom of an element has 6 electrons in the outermost shell. Can it be considered a noble gas if it accepts 2 electrons? Comment.
 8. "Noble gases such as Argon, Neon, Xenon are unreactive." Give reason in support of your answer.
 9. Calculate the number of protons, electrons and neutrons in Chlorine, Bromine, Iodine, Phosphorus, Silver and Gold. Also mention their atomic number and atomic mass (mass number).
 10. Is there any old thought about valency? If yes, what is the new concept of valency. Give reasons in support of your answer.

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